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4.2 For the reaction,  $2A + B \longrightarrow A_2B$ , the rate =  $k[A][B]^2$  with  $k = 2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$ . Calculate the initial rate of the reaction when  $[A] = 0.1 \text{ mol L}^{-1}$ ,  $[B] = 0.2 \text{ mol L}^{-1}$ . Calculate the rate of reaction after  $[A]$  is reduced to  $0.06 \text{ mol L}^{-1}$ . Sol. Initial rate of reaction =  $k[A][B]^2 = (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ s}^{-1})(0.1 \text{ mol L}^{-1})(0.2 \text{ mol L}^{-1})^2 = 8 \times 10^{-9} \text{ mol L}^{-1} \text{ s}^{-1}$ .

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In modern theory, it is now referred as coordination number of central metal atom or ion. 9.2  $\text{FeSO}_4$  solution mixed with  $(\text{NH}_4)_2\text{SO}_4$  solution in 1:1 molar ratio gives the test of  $\text{Fe}^{2+}$  ion but  $\text{CuSO}_4$  solution mixed with aqueous ammonia in 1:4 molar ratio does not give the test of  $\text{Cu}^{2+}$  ion. Explain why. Ans.  $\text{FeSO}_4$  solution mixed with  $(\text{NH}_4)_2\text{SO}_4$  solution in 1 : 1 molar ratio forms a double ...

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3.7 Define conductivity and molar conductivity for the solution of an electrolyte. Discuss their variation with concentration. Sol. Conductivity of a solution is defined as the conductance of a solution of 1 cm length and having 1 sq. cm as the area of cross-section. Molar conductivity of a solution at a dilution (V) is the conductance of all the ions produced from one mole of the electrolyte ...

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**12 Chemistry Solid State NCERT In Text Solution part 1**

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