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Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between the numbers of moles

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VIBRATIONS AND WAVES

Study Guide for Chapter 11

-Stoichiometry (Rough outline of the chapter, please use the book, notes & homework to study.) 11.1 Defining Stoichiometry Vocab • stoichiometry • mole ratio Concepts Using Balanced Equations • Number of Atoms • Number of Molecules • Number of Moles • Mass o

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Law of Conservation of Mass • Volume

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CHAPTER 11 SECTIONS 1 Defining
Stoichiometry 2 Stoichiometric
Calculations 3 Limiting Reactants 4
Percent Yield LaunchLAB What evidence
can you observe that a reaction has

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stopped? During a chemical reaction, reactants are consumed as new products form. In this lab, you will look for signs a chemical reaction has stopped. Steps in Stoichiometric Calculations

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15.2 CHAPTER 11: STOICHIOMETRY.
MOLE TO MOLE RATIO. When nitrogen and hydrogen gas are heated under the correct conditions, ammonia gas (NH_3) is formed. a. RXN: $1. \text{N}_2 + 3. \text{H}_2 \rightarrow 2. \text{NH}_3$. b. How many moles of nitrogen react with three moles of hydrogen? 1 mol N_2 3 mol H_2 1 mol N_2 . 3 mol H_2 . c.

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CHAPTER 11: STOICHIOMETRY

368 Chapter 11 • Stoichiometry Section
111.11.1 Objectives Describe the types
of relationships indicated by a balanced
chemical equation. State the mole ratios
from a balanced chemical equation.
Review Vocabulary reactant: the starting
substance in a chemical reaction New

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Vocabulary stoichiometry mole ratio
Defining Stoichiometry

Chapter 11: Stoichiometry

Chapter 11 Stoichiometry.
stoichiometry. mole ratio. excess
reactant. limiting reactant. The study of
quantitative relationships between the
amounts of.... In a balanced equation,

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the ratio between the numbers of moles.... A reactant that remains after a chemical reaction stops.

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homework to study.) 11.1 Defining
Stoichiometry Vocab • stoichiometry •
mole ratio Concepts Using Balanced
Equations • Number of Atoms • Number
of Molecules • Number of Moles • Mass o
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11.2 Stoichiometric Calculations
Concepts Mole-Mole ...

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for

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instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

Chapter 11.4: Stoichiometry -

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stoichiometry. mole ratio. excess

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reactant. limiting reactant. The study of quantitative relationships between the amounts of.... In a balanced equation, the ratio between the numbers of moles.... A reactant that remains after a chemical reaction stops.

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