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gas constant and the ratio of specific heats of oxygen. The units are: pressure N/m<sup>2</sup>, temperature K, area m<sup>2</sup> and mass flow rate kg/s. Assuming that the temperature of the oxygen in the bottle does not change with time, determine the time it takes to reduce the pressure to one half of its initial value.  $V=0.1 \text{ m}^3$

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$p_1 = 10 \text{ MPa}$   $T_1 = 293\text{K} = T_2$   $p_2 = 5 \text{ MPa}$   
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